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@ Mrs Collier HCo

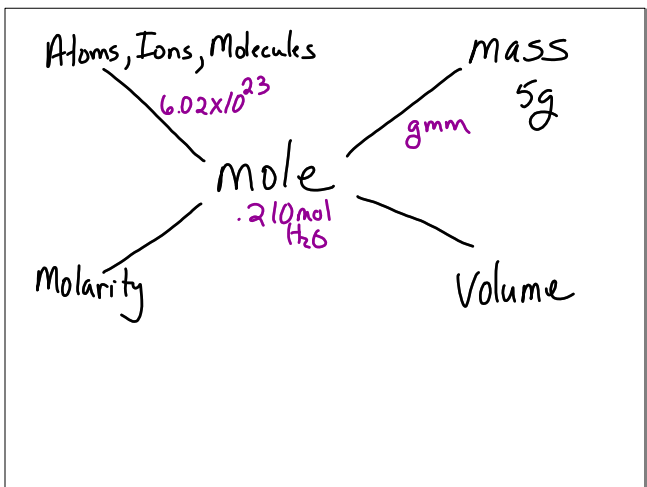
Aug 25-7:49 AM

$$\frac{16.15}{2.7} = \frac{5.98148}{6.0} + \frac{1.586}{2.31}$$

$$1,000,003.896$$

390

Aug 25-8:01 AM



Aug 25-8:03 AM

gmm = gram molecular mass
mm = molar mass

$$C = \frac{12.01 \text{ g}}{1 \text{ mol}} \quad 5 \text{ gC} \left| \frac{1 \text{ mol}}{12.01 \text{ g}} = .00416 \text{ mol} \right.$$

.004 mol

Aug 25-8:07 AM

H₂O

$$81 \text{ g H}_2\text{O} \left| \frac{1 \text{ mol}}{18.02 \text{ g}} = 4.5 \text{ mol H}_2\text{O} \right.$$

$$1.01(2) = 2.02$$

$$16.0(1) = 16.0$$

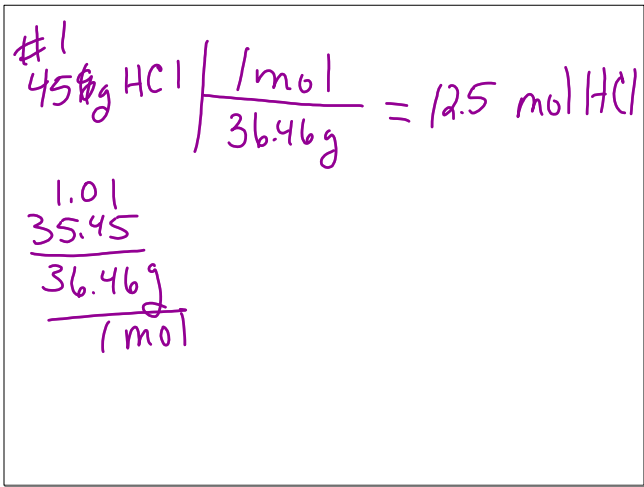
$$18.02 \text{ g} \left| \frac{1 \text{ mol}}{18.02 \text{ g}} = .21032 \text{ mol H}_2\text{O} \right.$$

Aug 25-8:11 AM

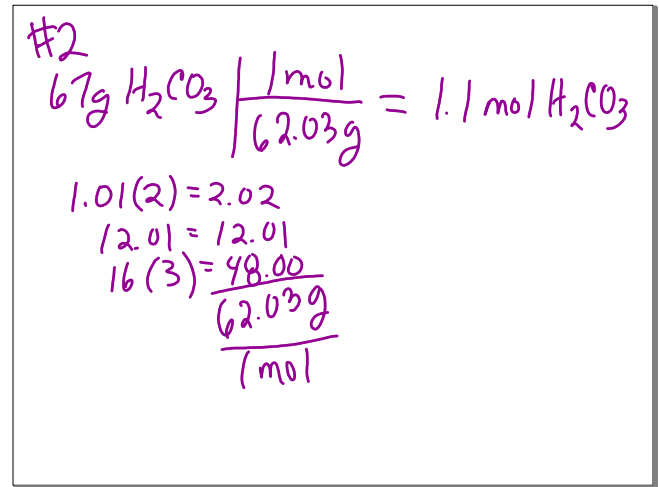
$$.210 \text{ mol H}_2\text{O} \left| \frac{6.02 \times 10^{23}}{1 \text{ mol}} = 210 \times 6.02 \text{ E } 23 \right.$$

$$.210 (6.02 \times 10^{23})$$

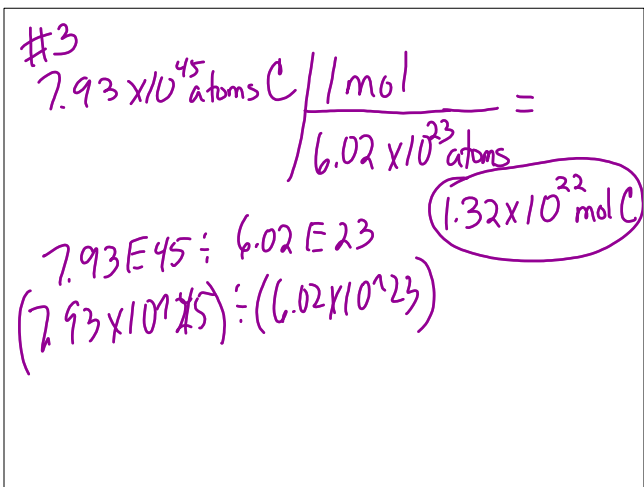
Aug 25-8:17 AM



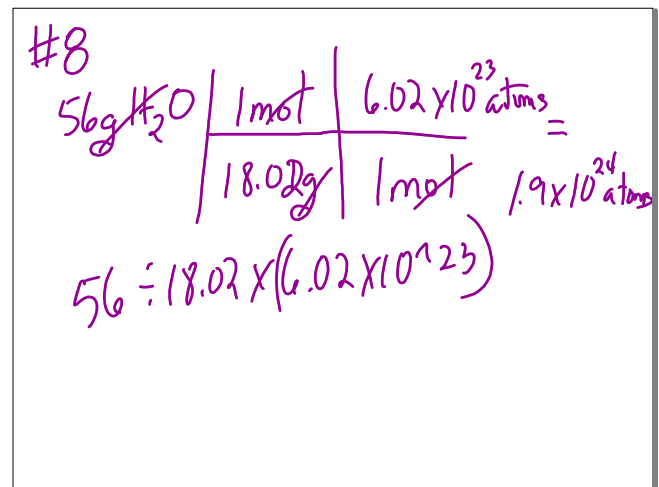
Aug 25-8:20 AM



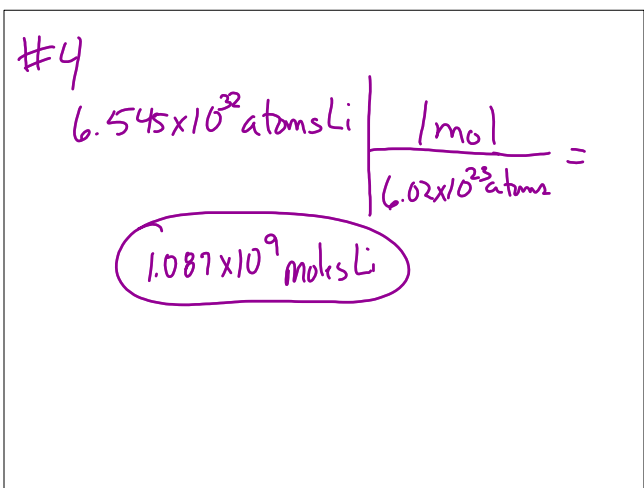
Aug 25-8:21 AM



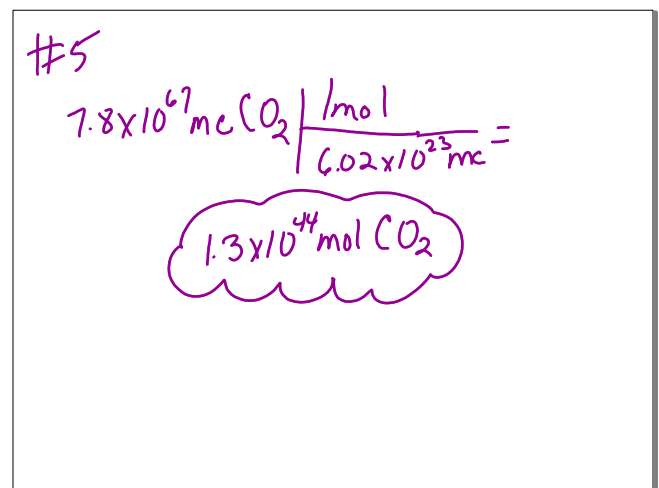
Aug 25-8:23 AM



Aug 25-8:27 AM



Aug 25-9:29 AM



Aug 25-9:31 AM

#6

$$56 \text{ mol H}_2\text{O} \left| \frac{6.02 \times 10^{23} \text{ mc}}{1 \text{ mol}} \right. =$$

$$\textcircled{3.4 \times 10^{25} \text{ mc H}_2\text{O}}$$

Aug 25-9:32 AM

#7

$$56 \text{ mol H}_2\text{O} \left| \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mol}} \right. =$$

$$3.4 \times 10^{25} \text{ atoms H}_2\text{O} \times 2 =$$

↑ 2 H's in H₂O

$$\textcircled{6.8 \times 10^{25} \text{ atoms H}}$$

Aug 25-9:32 AM

#9

$$89 \text{ g CaCl}_2 \left| \frac{1 \text{ mol}}{110.98 \text{ g}} \right. = .802 \text{ mol}$$

$$.802 \text{ mol} \left| \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mol}} \right. = 4.8 \times 10^{23} \text{ atoms}$$

↑ g mm

$$\times 2 \text{ Cl}$$

$$\textcircled{9.7 \times 10^{23} \text{ atoms}}$$

Aug 25-9:33 AM

#10

$$74 \text{ g C} \left| \frac{1 \text{ mol}}{12.0 \text{ g}} \right| \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mol}} =$$

$$\textcircled{3.7 \times 10^{24} \text{ atoms}}$$

Aug 25-9:34 AM