

12. Element J is the most non-metallic element in its group.
13. Elements C & N are from the same period and react to form CN .
14. The oxidation state (ion charge) of element O is 2^- in almost all compounds except when bonded to atoms of element Q.
15. Element Z is the most active metal in its period.
16. The atomic number of Z is 16 more than the atomic number of Y.
17. Element E reacts with element P to form EP_3 .
18. Element U is the smallest (size of the atoms) element in its period , but not in its group.
19. Element G has the highest ionization energy in its group.

The following Periodic Table shows only the Group A elements. Note that the atomic numbers jump from 20 to 31 in the fourth period. This is because elements 21 through 30, the transition elements, belong to the B Groups.

IA								VIII A
1 X		IIA	IIIA	IVA	VA	VIA	VIIA	2 S
3 Y	4 B	5 F	6 G	7 J	8 O	9 Q	10 U	
11 W	12 A	13 D	14 H	15 K	16 M	17 P	18 V	
19 Z	20 C	31 E	32 I	33 L	34 R	35 R	36 T	